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$\text{NH}_3$  2)  $2 \text{ KClO}_3 \rightleftharpoons 2 \text{ KCl} + 3 \text{ O}_2$  3)  $2$

$\text{NaCl} + 1 \text{ F}_2 \rightleftharpoons 2 \text{ NaF} + 1 \text{ Cl}_2$  4)  $2 \text{ H}_2 + 1$

$\text{O}_2 \rightleftharpoons 2 \text{ H}_2\text{O}$  5)  $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightleftharpoons 2$

$\text{H}_2\text{O} + 1 \text{ PbCl}_2$  6)  $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightleftharpoons$

$6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$  7)  $1 \text{ CH}_4 + 2 \text{ O}_2 \rightleftharpoons$

$1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$

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blank instead of writing "1" - in chemistry, they mean the same thing. 1)  $1 \text{ AlBr}_3 + 3 \text{ K} \rightarrow 3 \text{ KBr} + 1 \text{ Al}$  2)  $1 \text{ FeO} + 1 \text{ PdF}_2 \rightarrow 1 \text{ FeF}_2 + 1 \text{ PdO}$  3)  $1 \text{ P}_4 + 6 \text{ Br}_2 \rightarrow 4 \text{ PBr}_3$  4)  $2 \text{ LiCl} + 1 \text{ Br}_2 \rightarrow 2 \text{ LiBr} + 1 \text{ Cl}_2$  5)  $1 \text{ PbBr}_2 + 2 \text{ HCl} \rightarrow 2 \text{ HBr} + 1 \text{ PbCl}_2$  6)  $2 \text{ CoBr}_3 + 3 \text{ CaSO}_4 \dots$

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www.chemfiesta.com Balancing Equations Worksheet - Answers Note to students: It is acceptable to leave spaces blank when balancing equations - blank spaces are interpreted as containing the number "1".

1)  $1 \text{ Na}_3\text{PO}_4 + 3 \text{ KOH} \rightarrow 3 \text{ NaOH} + 1 \text{ K}_3\text{PO}_4$

2)  $1 \text{ MgF}_2 + 1 \text{ Li}_2\text{CO}_3 \rightarrow 1 \text{ MgCO}_3$

## Balancing Equations Worksheet - My Chemistry Class

Balancing Chemical Equations

### Balancing Chemical Equations

Answers for Predicting Products of Chemical Reactions For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be  $\text{CO}_2$  and  $\text{H}_2\text{O}$  (assuming a complete reaction).

a. Combustion:  $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$

b. Combustion:  $2\text{C}_4\text{H}_6 + 11\text{O}_2 \rightarrow 8\text{CO}_2 + 6\text{H}_2\text{O}$

c. Combustion:  $\text{C}_6\text{H}_{10}\text{O}_3 + 7\text{O}_2 \rightarrow 6\text{CO}_2 + 5\text{H}_2\text{O}$

1. Synthesis:  $\text{Mg} + \text{I}_2 \rightarrow \text{MgI}_2$  Note that Mg ...

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