

## Chapter 4 Atomic Structure Henry County School

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## **Chapter 4 Atomic Structure Henry**

Section 4.2 – Structure of an Atom. Protons, electrons, and neutrons are subatomic particles. A proton is a positively charged particle subatomic particle that is found in the nucleus of an atom. A proton has a +1 charge.

## **Chapter 4 Atomic Structure - Henry County School District**

4.2 The Structure of an Atom Protons, electrons, and neutrons are subatomic particles. • A proton is a positively charged subatomic particle that is found in the nucleus of an atom. It has a charge of 1+. • An electron is a negatively charged subatomic particle that is found in the space outside the nucleus. It has a charge of 1-. The mass of about

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Mass number vs Atomic Mass Atomic Mass- Atomic mass is the average of all the naturally occurring isotopes of that element. the decimal you see on the periodic table An average of all the isotopes of that element. Mass number- the number of protons and neutrons in the nucleus of a specific isotope The mass of that specific isotope. Atomic

## **Chapter 4 Atomic Structure - [schoolwires.henry.k12.ga.us](http://schoolwires.henry.k12.ga.us)**

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Section 4.1 Defining the Atom The Greek philosopher Democritus (460 B.C. - 370 B.C.) was among the first to suggest the existence of atoms (from the Greek word "atomos") He believed that atoms were indivisible and indestructible His ideas did agree with later scientific theory ...

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## **Chapter 4 Atomic Structure Review Sheet Answer Key**

Section 4.2 Structure of the Nuclear Atom One change to Dalton's atomic theory is that atoms are divisible into subatomic particles: Electrons, protons, and neutrons are examples of these fundamental particles There are many other types of particles, but we will study these three

## **Chapter 4 Atomic Structure - Henry County School District ...**

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## **Chapter 4 Atomic Structure Section Review Answers**

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Chapter 4 Atomic Structure Section 4.3 Modern Atomic Theory (pages 113–118) This section focuses on the arrangement and behavior of electrons in atoms. Reading Strategy (page 113) Sequencing After you read, complete the description in the flow chart below of how the gain or loss of energy affects electrons in atoms. For

### **Chapter 4 Atomic Structure Section 4.3 Modern Atomic Theory**

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### **Chapter 4 Atomic Structure Flashcards | Quizlet**

Notes Chapter 4: The Structure of the Atom. 4.1 Early Theories of Matter. A. The Philosophers - believed matter was made of earth, water, air, and fire. 1. Democritus (460-370 BC) was the first to propose that matter was made up of tiny \_\_\_\_\_ he called atomos,

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which could not be further \_\_\_\_\_. a.

### **Notes Chapter 4: The Structure of the Atom**

Chapter 4 (Atomic Structure) Quiz study guide by jgreenzaid includes 37 questions covering vocabulary, terms and more. Quizlet flashcards, activities and games help you improve your grades.

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alpha particle = helium; 4 top, 2 bottom  
beta particle = electron; 0 top, -1 bottom  
gamma = 0

### **Chapter 4: The Structure of an Atom Flashcards | Quizlet**

Since the atomic number of sodium atom is 11, it has 11 electrons. A positively charged sodium ion ( $\text{Na}^+$ ) is formed by the removal of one electron from a sodium atom. So, a sodium ion has  $11 - 1 = 10$  electrons in it. Thus, electronic distribution of

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sodium ion will be 2, 8.

## **Chapter 4: Structure of Atom - Olympiad Success**

Chapter 1 & 2; AP Chem Virtual Bonding Lab; Honors & General Chemistry. First 9 Weeks; Second 9 Weeks; Online Chemistry Resources; Standards; Virtual Chemistry Labs; Flipping Chemistry . Introduction to Stoichiometry; Stoichiometry - Mole Ratio; Stoichiometry - Multi Step Conversions; Unit 4 Naming, Bonding & Intro to calculations recovery

## **Science: Raines, Charlena / First 9 Weeks**

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## **Chemistry - Chp 4 - Atomic Structure - PowerPoint**

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41. Measuring Atomic Mass □ Instead of grams, the unit we use is the Atomic Mass Unit (amu) □ It is defined as one-twelfth the mass of a carbon-12 atom. □ Carbon-12 chosen because of its isotope purity. □ Each isotope has its own atomic mass,...

### **Chemistry - Chp 4 - Atomic Structure - PowerPoint**

date: \_\_\_\_\_ CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure Matching. A. Bohr B. Democritus C. Rutherford D. Dalton E. Thomson F. Schrodinger \_\_\_\_\_ 1. Greek thinker; called nature's basic particle an atom, based on the Greek word "atomos" which means "indivisible". Did not have evidence that atoms existed.

### **CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure**

Chapter 4 Atomic Structure Section 4.1 Studying Atoms (pages 100-105) This section discusses the development of atomic



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models. Reading Strategy(page 100) Summarizing As you read, complete the table about atomic models. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of

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